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Methyl 1-[4-[(S)-2-(methoxycarbonyl)-pyrrolidin-1-yl]-3,6-dioxocyclohexa-1,4-dien-1-yl]pyrrolidine-2-carboxylate

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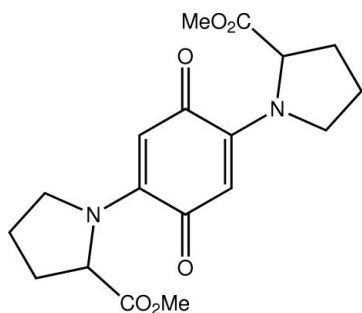
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Key indicators: single-crystal X-ray study; $T = 120$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.038; wR factor = 0.136; data-to-parameter ratio = 9.4.

The complete molecule of the title diproline ester quinone, $\text{C}_{18}\text{H}_{22}\text{N}_2\text{O}_6$, is generated by a crystallographic twofold axis, which passes through the centre of the benzene ring. Both $-\text{CO}_2\text{Me}$ groups are orientated to the same side of the benzene ring, with the carbonyl groups pointing roughly towards each other. The conformation of the proline residue is an envelope. In the crystal, a three-dimensional network is sustained by $\text{C}-\text{H}\cdots\text{O}$ interactions involving both the quinone and carbonyl O atoms.

Related literature

For the oxidative nucleophilic addition of amines to quinones to form aminoquinones, see: Lyons & Thomson (1953). For background to mitomycin anticancer drugs, see: Tomasz (1995). For additional geometric analysis, see: Cremer & Pople (1975).



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Experimental

Crystal data

$\text{C}_{18}\text{H}_{22}\text{N}_2\text{O}_6$
 $M_r = 362.38$
 Monoclinic, $C2$
 $a = 11.4728$ (5) Å
 $b = 7.1556$ (4) Å
 $c = 11.7882$ (7) Å
 $\beta = 111.230$ (3)°
 $V = 902.07$ (8) Å³
 $Z = 2$
 Mo $K\alpha$ radiation
 $\mu = 0.10$ mm⁻¹
 $T = 120$ K
 $0.24 \times 0.12 \times 0.08$ mm

Data collection

Nonius KappaCCD area-detector diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 2007)
 $T_{\min} = 0.896$, $T_{\max} = 1.000$
 6840 measured reflections
 1114 independent reflections
 1008 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.034$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.038$
 $wR(F^2) = 0.136$
 $S = 1.23$
 1114 reflections
 119 parameters
 1 restraint
 H-atom parameters constrained
 $\Delta\rho_{\max} = 0.64$ e Å⁻³
 $\Delta\rho_{\min} = -0.63$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C3}-\text{H3}\cdots\text{O2}^{\text{i}}$	0.95	2.56	3.400 (3)	147
$\text{C5}-\text{H5b}\cdots\text{O1}^{\text{ii}}$	0.99	2.54	3.407 (3)	146
$\text{C9}-\text{H9b}\cdots\text{O1}^{\text{iii}}$	0.98	2.38	3.186 (4)	139

Symmetry codes: (i) $-x + \frac{1}{2}, y - \frac{1}{2}, -z$; (ii) $x + \frac{1}{2}, y - \frac{1}{2}, z$; (iii) $-x + \frac{1}{2}, y + \frac{1}{2}, -z + 1$.

Data collection: *COLLECT* (Hooft, 1998); cell refinement: *DENZO* (Otwinowski & Minor, 1997) and *COLLECT*; data reduction: *DENZO* and *COLLECT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997) and *DIAMOND* (Brandenburg, 2006); software used to prepare material for publication: *publCIF* (Westrip, 2010).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB5550).

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supplementary materials

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Methyl 1-{4-[(*S*)-2-(methoxycarbonyl)pyrrolidin-1-yl]-3,6-dioxocyclohexa-1,4-dien-1-yl}pyrrolidine-2-carboxylate

S. J. Garden, J. M. S. Skakle, E. R. T. Tiekink and J. L. Wardell

Comment

Oxidative nucleophilic addition of amines to quinones results in the formation of aminoquinone products (Lyons & Thomson, 1953). As part of a study into concise methodology for the synthesis of heterocyclic systems, we envisaged that oxidative addition of α -amino acid derivatives to benzoquinone could yield a suitably functionalized precursor for cyclization to yield pyrroloindole quinones, a structural motif present in the mitomycin anticancer drugs (Tomasz, 1995). The title di-proline ester quinone, (I), was synthesized in this context.

The molecule of (I), Fig. 1, exists about a crystallographic 2-fold axis of symmetry passing through the centre of the benzene ring. This has the result that the two $-\text{CO}_2\text{Me}$ groups are orientated to the same side of the benzene ring. The carbonyl groups are tucked in under the benzene ring. The conformation of the proline residue is an envelope with the C5 atom lying above the plane through the remaining atoms. The conformational descriptors (Cremer & Pople, 1975) are $Q(2) = 0.348(3) \text{ \AA}$ and $\varphi(2) = 79.3(4)^\circ$.

The crystal packing features C—H \cdots O contacts, Table 1. The quinone-O1 atom accepts two such interactions, one from a methylene-H and the other from a methyl-H, whereas the carbonyl-O2 accepts a quinone-H. The C—H \cdots O interactions combine to give a 3-D network, Fig. 2.

Experimental

Proline methyl ester hydrochloride (1.63 g, 9.8 mmol) and KOAc (1.07 g, 10.9 mmol) were mixed in MeOH (20 ml). Excess benzoquinone (1.00 g, 9.3 mmol) was added to the solution resulting in a deep-red coloured reaction. The reaction was stirred for an hour then all the volatiles were removed under reduced pressure. The crude product was solubilized in EtOAc and filtered through a plug of silica using EtOAc as eluent. The red coloured fraction was evaporated under reduced pressure and the product chromatographed on a column of silica eluting with CH_2Cl_2 (removed excess benzoquinone) followed by a $\text{CH}_2\text{Cl}_2/\text{EtOAc}$ gradient (9:1 V/V to 4:1 V/V). Evaporation of the product containing fractions and recrystallization from MeOH gave 0.266 g of dark-red prisms of (I) (15% yield), m.pt. 463–465 K.

^1H (CDCl_3): δ 1.95 [2H, m]; 2.19 [2H, m]; 3.39 [1H, m]; 3.48 [1H, m]; 3.73 [3H, s]; 5.07 [1H, bs]; 5.37 [1H, bs] p.p.m.
 ^{13}C (CDCl_3): δ 22.0; 31.6; 51.3; 52.4; 62.9; 101.5; 148.7; 172.8; 181.2 p.p.m. IR (cm^{-1}): 3066, 2982, 2955, 2927, 2881, 1744, 1625, 1555, 1432, 1349, 1274, 1207, 1166, 957, 833, 794. Mass (a.m.u.) abundance %: 362 (79), 331 (10), 303 (52), 276 (100), 261 (63), 243 (25), 235 (37), 217 (46), 176 (25), 122 (37).

Refinement

The C-bound H atoms were geometrically placed (C–H = 0.95–1.00 Å) and refined as riding with $U_{\text{iso}}(\text{H}) = 1.2\text{--}1.5U_{\text{eq}}(\text{C})$. The maximum and minimum residual electron density peaks of 0.64 and 0.63 e Å⁻³, respectively, were located 1.59 Å and 0.85 Å from the H5b and C8 atoms, respectively. In the absence of significant anomalous scattering effects, 750 Friedel pairs were averaged in the final refinement. However, the absolute configuration was assigned on the basis of the chirality of the *L*-proline starting material.

Figures

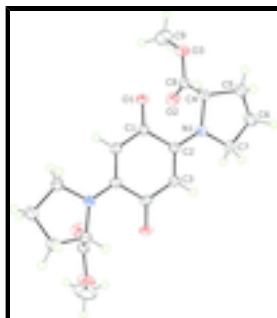


Fig. 1. The molecular structure of (I) showing displacement ellipsoids at the 50% probability level. Unlabelled atoms are generated by $(-x, y, -z)$.

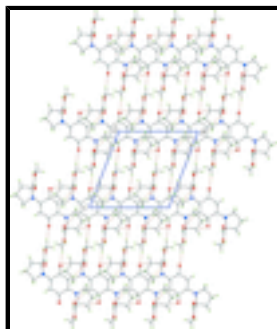


Fig. 2. A view in projection down the *b* axis of the unit-cell contents (I) showing the C–H···O contacts as orange dashed lines.

Methyl 1-{4-[(*S*)-2-(methoxycarbonyl)pyrrolidin-1-yl]-3,6-dioxocyclohexa-1,4-dien-1-yl}pyrrolidine-2-carboxylate

Crystal data

C₁₈H₂₂N₂O₆

$M_r = 362.38$

Monoclinic, *C*2

Hall symbol: C 2y

$a = 11.4728$ (5) Å

$b = 7.1556$ (4) Å

$c = 11.7882$ (7) Å

$\beta = 111.230$ (3)°

$V = 902.07$ (8) Å³

$Z = 2$

$F(000) = 384$

$D_x = 1.334$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 1107 reflections

$\theta = 2.9\text{--}27.5^\circ$

$\mu = 0.10$ mm⁻¹

$T = 120$ K

Prism, dark-red

$0.24 \times 0.12 \times 0.08$ mm

Data collection

Nonius KappaCCD area-detector diffractometer	1114 independent reflections
Radiation source: Enraf Nonius FR591 rotating anode	1008 reflections with $I > 2\sigma(I)$
10 cm confocal mirrors	$R_{\text{int}} = 0.034$
Detector resolution: 9.091 pixels mm^{-1}	$\theta_{\text{max}} = 27.5^\circ$, $\theta_{\text{min}} = 3.4^\circ$
φ and ω scans	$h = -14 \rightarrow 14$
Absorption correction: multi-scan (SADABS; Sheldrick, 2007)	$k = -8 \rightarrow 9$
$T_{\text{min}} = 0.896$, $T_{\text{max}} = 1.000$	$l = -15 \rightarrow 15$
6840 measured reflections	

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.038$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.136$	H-atom parameters constrained
$S = 1.23$	$w = 1/[\sigma^2(F_o^2) + (0.0929P)^2]$
1114 reflections	where $P = (F_o^2 + 2F_c^2)/3$
119 parameters	$(\Delta/\sigma)_{\text{max}} < 0.001$
1 restraint	$\Delta\rho_{\text{max}} = 0.64 \text{ e } \text{\AA}^{-3}$
	$\Delta\rho_{\text{min}} = -0.63 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.08333 (15)	0.9788 (3)	0.24336 (14)	0.0259 (4)
O2	0.27854 (17)	1.2823 (3)	0.25151 (17)	0.0334 (5)
O3	0.34432 (18)	1.1322 (4)	0.43176 (16)	0.0417 (6)
N1	0.25815 (18)	0.9487 (3)	0.12794 (16)	0.0218 (5)

supplementary materials

C1	0.0437 (2)	0.9680 (4)	0.1309 (2)	0.0201 (5)
C2	0.1346 (2)	0.9582 (4)	0.06419 (19)	0.0195 (5)
C3	0.0870 (2)	0.9607 (4)	-0.06108 (19)	0.0211 (5)
H3	0.1438	0.9574	-0.1029	0.025*
C4	0.3199 (2)	0.9490 (4)	0.2607 (2)	0.0251 (5)
H4	0.2809	0.8534	0.2977	0.030*
C5	0.4542 (2)	0.8929 (4)	0.2793 (2)	0.0297 (6)
H5A	0.5139	0.9458	0.3559	0.036*
H5B	0.4635	0.7552	0.2815	0.036*
C6	0.4758 (2)	0.9750 (5)	0.1691 (2)	0.0291 (6)
H6A	0.5411	0.9044	0.1508	0.035*
H6B	0.5013	1.1078	0.1830	0.035*
C7	0.3489 (2)	0.9556 (4)	0.0660 (2)	0.0236 (5)
H7A	0.3324	1.0638	0.0101	0.028*
H7B	0.3454	0.8397	0.0190	0.028*
C8	0.3099 (2)	1.1413 (4)	0.3104 (2)	0.0261 (6)
C9	0.3358 (4)	1.3085 (7)	0.4894 (3)	0.0656 (12)
H9A	0.3767	1.4069	0.4595	0.098*
H9B	0.3772	1.2966	0.5778	0.098*
H9C	0.2476	1.3409	0.4698	0.098*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0265 (8)	0.0356 (11)	0.0146 (7)	-0.0001 (8)	0.0063 (6)	0.0006 (8)
O2	0.0369 (10)	0.0308 (11)	0.0359 (10)	-0.0014 (9)	0.0174 (8)	-0.0048 (9)
O3	0.0396 (11)	0.0594 (14)	0.0204 (9)	0.0108 (11)	0.0040 (8)	-0.0119 (10)
N1	0.0203 (9)	0.0265 (11)	0.0177 (9)	0.0016 (9)	0.0057 (7)	-0.0024 (9)
C1	0.0233 (10)	0.0196 (11)	0.0180 (9)	0.0033 (10)	0.0082 (8)	0.0030 (10)
C2	0.0225 (10)	0.0185 (11)	0.0174 (10)	0.0002 (11)	0.0071 (8)	-0.0006 (10)
C3	0.0232 (10)	0.0234 (11)	0.0186 (10)	-0.0013 (10)	0.0097 (8)	-0.0033 (10)
C4	0.0224 (11)	0.0328 (14)	0.0175 (10)	0.0019 (12)	0.0041 (8)	-0.0005 (11)
C5	0.0265 (12)	0.0330 (14)	0.0256 (12)	0.0064 (11)	0.0046 (9)	-0.0003 (11)
C6	0.0219 (11)	0.0326 (14)	0.0309 (12)	-0.0004 (11)	0.0072 (9)	-0.0011 (12)
C7	0.0219 (11)	0.0255 (12)	0.0247 (10)	0.0023 (11)	0.0100 (8)	-0.0003 (11)
C8	0.0175 (11)	0.0380 (15)	0.0215 (11)	0.0021 (11)	0.0055 (8)	-0.0045 (11)
C9	0.064 (2)	0.088 (3)	0.0364 (16)	0.017 (2)	0.0078 (15)	-0.034 (2)

Geometric parameters (\AA , $^\circ$)

O1—C1	1.238 (3)	C4—C5	1.530 (3)
O2—C8	1.203 (4)	C4—H4	1.0000
O3—C8	1.341 (3)	C5—C6	1.525 (4)
O3—C9	1.453 (5)	C5—H5A	0.9900
N1—C2	1.345 (3)	C5—H5B	0.9900
N1—C4	1.466 (3)	C6—C7	1.528 (3)
N1—C7	1.473 (3)	C6—H6A	0.9900
C1—C3 ⁱ	1.426 (3)	C6—H6B	0.9900

C1—C2	1.518 (3)	C7—H7A	0.9900
C2—C3	1.377 (3)	C7—H7B	0.9900
C3—C1 ⁱ	1.426 (3)	C9—H9A	0.9800
C3—H3	0.9500	C9—H9B	0.9800
C4—C8	1.516 (4)	C9—H9C	0.9800
C8—O3—C9	114.4 (3)	C4—C5—H5B	110.9
C2—N1—C4	126.92 (18)	H5A—C5—H5B	109.0
C2—N1—C7	120.87 (18)	C5—C6—C7	103.91 (19)
C4—N1—C7	111.96 (18)	C5—C6—H6A	111.0
O1—C1—C3 ⁱ	121.5 (2)	C7—C6—H6A	111.0
O1—C1—C2	120.2 (2)	C5—C6—H6B	111.0
C3 ⁱ —C1—C2	118.3 (2)	C7—C6—H6B	111.0
N1—C2—C3	121.90 (19)	H6A—C6—H6B	109.0
N1—C2—C1	119.72 (19)	N1—C7—C6	104.46 (19)
C3—C2—C1	118.4 (2)	N1—C7—H7A	110.9
C2—C3—C1 ⁱ	123.07 (19)	C6—C7—H7A	110.9
C2—C3—H3	118.5	N1—C7—H7B	110.9
C1 ⁱ —C3—H3	118.5	C6—C7—H7B	110.9
N1—C4—C8	109.8 (2)	H7A—C7—H7B	108.9
N1—C4—C5	102.97 (18)	O2—C8—O3	124.5 (3)
C8—C4—C5	113.3 (2)	O2—C8—C4	126.0 (2)
N1—C4—H4	110.2	O3—C8—C4	109.5 (2)
C8—C4—H4	110.2	O3—C9—H9A	109.5
C5—C4—H4	110.2	O3—C9—H9B	109.5
C6—C5—C4	104.1 (2)	H9A—C9—H9B	109.5
C6—C5—H5A	110.9	O3—C9—H9C	109.5
C4—C5—H5A	110.9	H9A—C9—H9C	109.5
C6—C5—H5B	110.9	H9B—C9—H9C	109.5
C4—N1—C2—C3	178.7 (3)	C7—N1—C4—C5	-17.7 (3)
C7—N1—C2—C3	4.9 (4)	N1—C4—C5—C6	32.4 (3)
C4—N1—C2—C1	-0.9 (4)	C8—C4—C5—C6	-86.2 (2)
C7—N1—C2—C1	-174.7 (2)	C4—C5—C6—C7	-35.3 (3)
O1—C1—C2—N1	4.5 (4)	C2—N1—C7—C6	170.6 (2)
C3 ⁱ —C1—C2—N1	-174.4 (2)	C4—N1—C7—C6	-4.1 (3)
O1—C1—C2—C3	-175.1 (2)	C5—C6—C7—N1	24.3 (3)
C3 ⁱ —C1—C2—C3	6.1 (3)	C9—O3—C8—O2	2.6 (4)
N1—C2—C3—C1 ⁱ	179.1 (3)	C9—O3—C8—C4	-178.7 (2)
C1—C2—C3—C1 ⁱ	-1.4 (3)	N1—C4—C8—O2	-13.2 (3)
C2—N1—C4—C8	-71.0 (3)	C5—C4—C8—O2	101.3 (3)
C7—N1—C4—C8	103.3 (2)	N1—C4—C8—O3	168.10 (19)
C2—N1—C4—C5	168.0 (2)	C5—C4—C8—O3	-77.4 (2)

Symmetry codes: (i) $-x, y, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C3—H3 \cdots O2 ⁱⁱ	0.95	2.56	3.400 (3)	147

supplementary materials

C5—H5b···O1 ⁱⁱⁱ	0.99	2.54	3.407 (3)	146
C9—H9b···O1 ^{iv}	0.98	2.38	3.186 (4)	139

Symmetry codes: (ii) $-x+1/2, y-1/2, -z$; (iii) $x+1/2, y-1/2, z$; (iv) $-x+1/2, y+1/2, -z+1$.

Fig. 1

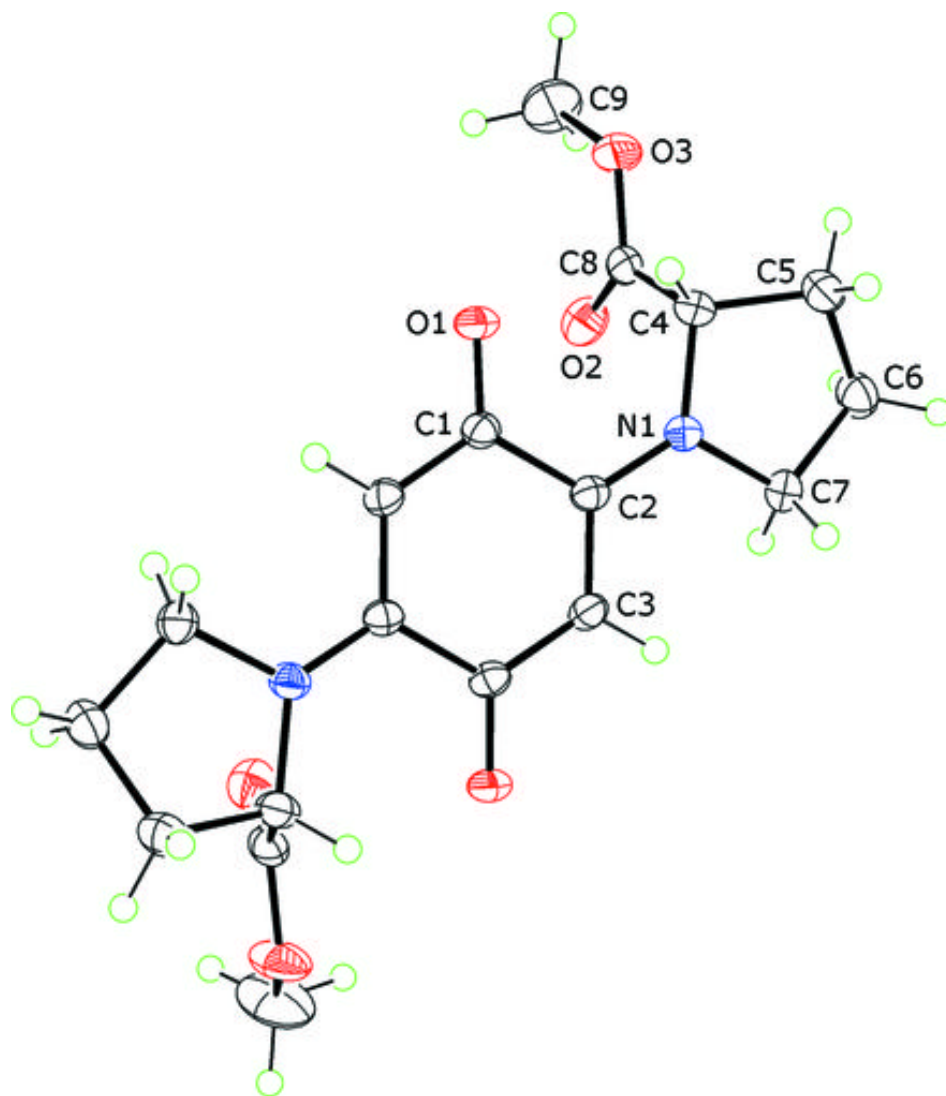


Fig. 2

